



100 Stoichiometry Questions

1. If 5 moles of hydrogen gas react with 2 moles of oxygen gas to form water, what is the limiting reactant?

Ans: Oxygen

Video explanation: <https://youtu.be/RD7kRfUX5Os>

2. If 4 g of hydrogen gas reacts with 2 g of oxygen gas to form water, what is the limiting reactant?

Ans: Oxygen

Video explanation: https://youtu.be/_kkw8TW-Mvw

3. How many moles of water are formed when 5 moles of hydrogen gas react with 2 moles of oxygen gas?

Ans: 4 moles of water

Video explanation: <https://youtu.be/sk6LI2GiWHg>

4. A reaction requires 3 moles of hydrogen gas. How many grams of aluminum, reacting with excess hydrochloric acid according to the equation: $2 \text{Al} + 6 \text{HCl} \rightarrow 2 \text{AlCl}_3 + 3 \text{H}_2$, are needed to produce the necessary hydrogen? [Al=27, Cl=35.5, H=1]

Ans: 54 g of Al

Video explanation: <https://youtu.be/Xsc7p3oqTaw>

5. Given the reaction: $4 \text{NH}_3 + 5 \text{O}_2 \rightarrow 4 \text{NO} + 6 \text{H}_2\text{O}$, if 10 moles of ammonia react, how many moles of oxygen are required?

Ans: 12.5 moles of O_2

Video explanation: <https://youtu.be/nyyONI1R4H4>

6. Given the reaction: $4 \text{NH}_3 + 5 \text{O}_2 \rightarrow 4 \text{NO} + 6 \text{H}_2\text{O}$, if 10 moles of ammonia react, how many moles of water are produced?

Ans: 15 moles of H_2O

Video explanation: <https://youtu.be/BIQctRscoAc>

7. If 2.5 moles of methane (CH_4) react with excess oxygen according to the equation:

$\text{CH}_4 + 2 \text{O}_2 \rightarrow \text{CO}_2 + 2 \text{H}_2\text{O}$, what mass of water is produced? [C=12, O=16, H=1]

Ans: 90 g of H_2O

Video explanation: <https://youtu.be/AsVwQhFImBY>

8. If 25 grams of methane (CH_4) react with excess oxygen according to the equation:

$\text{CH}_4 + 2 \text{O}_2 \rightarrow \text{CO}_2 + 2 \text{H}_2\text{O}$, what mass of water is produced? [C=12, O=16, H=1]

Ans: 56.25 g of H_2O

Video explanation: https://youtu.be/-5_09rYSryk

9. How many moles of oxygen are required to completely react with 15 moles of propane (C_3H_8) according to the balanced equation: $\text{C}_3\text{H}_8 + 5 \text{O}_2 \rightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O}$? [C=12, H=1, O=16]

Ans: 75 moles of O_2

Video explanation: <https://youtu.be/T4Ypkd77nu8>

10. A reaction produces 18 moles of product according to the balanced equation:

$3 \text{A} + 2 \text{B} \rightarrow 6 \text{C}$. If A is the limiting reactant, how many moles of B reacted?

Ans: 6 moles of B

Video explanation: <https://youtu.be/CKKqMHtVheM>

11. Given the reaction: $2 \text{K} + 2 \text{H}_2\text{O} \rightarrow 2 \text{KOH} + \text{H}_2$, if 10 g of potassium (K) react, how many grams of hydrogen (H_2) are produced? [K=39, H=1, O=16]

Ans: 0.26 g of H_2

Video explanation: <https://youtu.be/JWizM4ZJDj0>

12. How many grams of calcium oxide are produced when 50 g of calcium carbonate (CaCO_3) decomposes? The balanced equation is: $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$. [Ca=40, C=12, O=16]

Ans: 28 g of CaO

Video explanation: <https://youtu.be/TB8FoDkbJR0>

13. If 30 grams of nitrogen gas (N_2) react with excess hydrogen gas (H_2) according to the equation:

$\text{N}_2 + 3 \text{H}_2 \rightarrow 2 \text{NH}_3$, how many grams of ammonia (NH_3) are produced? [N=14, H=1]

Ans: 36 g NH_3

Video explanation: <https://youtu.be/4ZVT-KmOOHE>

14. $\text{N}_2 + 3 \text{H}_2 \rightarrow 2 \text{NH}_3$, how many litres of ammonia are formed from the complete reaction of 0.225 L of hydrogen?

Ans: 0.15 L of NH_3

Video explanation: https://youtu.be/l-_5KliZJ_E

15. How many atoms of Ne are there in 0.55 moles of Ne? [Avogadro's no= 6.022×10^{23}]

Ans: 3.3×10^{23} atoms

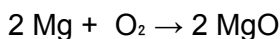
Video explanation: <https://youtu.be/3knJjDVrb64>

16. How many H atoms are present in 50 molecules of C₄H₄S₂?

Ans: 200 atoms of H

Video explanation: <https://youtu.be/FiYkJVA9tmc>

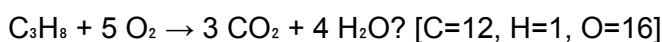
17. How many moles of O₂ are needed for reaction with 4 moles of Mg?



Ans: 2 moles of O₂

Video explanation: <https://youtu.be/mAoDCNnQExc>

18. How many moles of water are produced on combustion of propane with 2.5 moles of O₂ ?



Ans: 2 moles of water

Video explanation: https://youtu.be/PfKi_2OL7nk

19. How many moles are present in 1.5 g of CaCl₂ [Ca=40, Cl=35.5]?

Ans: 0.014 moles of CaCl₂

Video explanation: <https://youtu.be/mrmrDBFfDXE>

20. How many moles are present in 15 mL of 0.1 M AgNO₃ solution [Ag=108, N=14, O=16]?

Ans: 0.002 moles

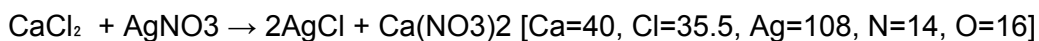
Video explanation: <https://youtu.be/-Tp15S-eCGI>

21. If 1.5 g of CaCl₂ and 15 mL of 0.1 M AgNO₃ are reacted, what is the limiting reagent? $\text{CaCl}_2 + \text{AgNO}_3 \rightarrow 2\text{AgCl} + \text{Ca}(\text{NO}_3)_2$ [Ca=40, Cl=35.5, Ag=108, N=14, O=16]

Ans: AgNO₃

Video explanation: <https://youtu.be/ul4Rd284H-Q>

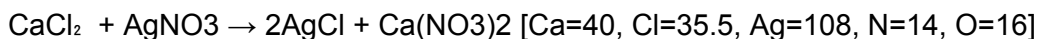
22. If 1.5 g of CaCl₂ and 15 mL of 0.1 M AgNO₃ are reacted, calculate the quantity of the excess reagent at the end of the reaction in moles?



Ans: 0.01 moles CaCl₂

Video explanation: <https://youtu.be/019WXE5Cylo>

23. If 1.5 g of CaCl₂ and 15 mL of 0.1 M AgNO₃ are reacted, calculate the amount of AgCl produced in moles?



Ans: 0.002 moles AgCl

Video explanation: <https://youtu.be/tLnDcdl3FuU>

24. How many moles of MgCl_2 are produced by reacting 150 g of MgCO_3 with excess HCl. [Mg=24, C=12, O=16, H=1]

Ans: 1.786 mol MgCl_2

Video explanation: <https://youtu.be/BOovfAyCpRM>

25. Calculate the mass of anhydrous Na_2CO_3 present in 500 mL of 0.1 M Na_2CO_3 solution.

Ans: 5 g

Video explanation: <https://youtu.be/4onQghdeQ0Q>

26. A metal X reacts with oxygen gas to form X_2O_3 . How many moles of metal react with 20 g of oxygen?

Ans: 0.83 moles of X

Video explanation: https://youtu.be/7HK3Evch_8A

27. How many moles of CO_2 are formed from the combustion of 5 g of ethane?

Ans: 0.3 moles of CO_2

Video explanation: <https://youtu.be/UhX7hkepIJQ>

28. Calculate the mass of ZnO produced from the decomposition of 50 g ZnCO_3 . [Zn=65, C=12, O=16].

Ans: 32 g

Video explanation: <https://youtu.be/fvdd5HJPUVg>

29. Calculate the volume of steam produced at s.t.p when 18 g of propyne is combusted.

Ans: 20 L

Video explanation:

30. 0.2 g of hydrogen gas combines with 3.2 g of oxygen gas. If the relative molecular mass of the product is 34, what is its molecular formula?

Ans: H_2O_2

Video explanation: <https://youtu.be/NNy0cp3dMV8>

31. Write an equation of reaction between a divalent metal X and HCl.

Ans: $\text{X(s)} + 2\text{HCl(aq)} \rightarrow \text{XCl}_2\text{(s)} + \text{H}_2\text{(g)}$

Video explanation: <https://youtu.be/lBemzL2sQL4>

32. A divalent metal X reacts with limited HCl to produce 600 mL of hydrogen gas at s.t.p. Calculate the mass of the metal X [molar vol = 22.4 L, X=24, H = 1, Cl=35.5]

Ans: 0.643 g of X

Video explanation: <https://youtu.be/13j-EAODY0>

33. How many moles of HCl is necessary to neutralise 10 mL of 0.5 M K_2CO_3 solution [H=1, Cl=35.5, K=39, C=12, O=16]?

Ans: 0.01 moles of HCl

Video explanation: <https://youtu.be/5vH0UNixYUA>

34. Calculate the mass of HCl left when a small piece of zinc metal (0.5 g) reacts with excess HCl. [Zn=65, H=1, Cl=35.5]

Ans: $(36.5 I - 0.5615)$ g where I is the initial moles of HCl

Video explanation: <https://youtu.be/wsn1gkKqKCU>

35. How many moles of hydrogen ions are present in a solution containing 3 moles of H_2SO_4 [H=1, S=32, O=16]?

Ans: 6 moles of hydrogen ions

Video explanation: <https://youtu.be/-SihXg8Frk4>

36. Calculate the mass of the Chloride ion present in 0.3 g of sodium chloride solution [Na=23, Cl=35.5]

Ans: 0.182 g

Video explanation: <https://youtu.be/QkGr0PWUcik>

37. What is the mole ratio of N_2 to H_2 in the stoichiometric equation for the production of NH_3 .

Ans: 1:3

Video explanation: <https://youtu.be/zNDUQodzDIk>

38. Calculate the number of ions present in 2 moles of potassium ion.

Ans: $1.2044E34$

Video explanation: <https://youtu.be/9F-1duf0YQs>

39. Calculate the mass of O_2 produced from the decomposition of 35 g $KClO_3$ [K=39, Cl=35.5, O=16].

Ans: 14 g

Video explanation: <https://youtu.be/2lO0yKFQa-c>

40. 40 g of a KCl and $KClO_3$ mixture were heated to produce a residue weighing 29 g. What is the percentage of KCl in the mixture [K=39, Cl=35.5, O=16] ?

Ans: 29.8%

Video explanation: <https://youtu.be/jltn4KeE4WE>

41. Calculate the mass of H_2SO_4 present in 24×10^{23} molecules of H_2SO_4

Ans: 391 g

Video explanation: <https://youtu.be/M2et8OvNLPc>

42. Calculate the number of ions in 1.2 g of Magnesium metal ($\text{Mg}=24$)

Ans: 3.011×10^{22}

Video explanation: <https://youtu.be/zSistYLGm0w>

43. Calculate the mass of lead (ii) sulphate produced when 0.5 moles of sulphuric acid reacts with excess lead (ii) chloride.

Ans: 151.5 g

Video explanation: <https://youtu.be/A5gw0s1qdmA>

44. Calculate the volume of ammonia gas produced at s.t.p when 0.3 moles of hydrogen gas reacts with excess nitrogen gas

Ans: 4.48 L

Video explanation: <https://youtu.be/t2rCHU3jkOY>

45. What mass of CaO is formed from completely burning 2 g of Ca in oxygen [$\text{Ca}=40$, $\text{O}=16$]?

Ans: 2.8 g

Video explanation: <https://youtu.be/fwj-gntv6So>

46. 6L Chlorine gas at s.t.p contains how many grams of chlorine.

Ans: 19.017 g

Video explanation: <https://youtu.be/pXDvxmwhfA4>

47. At s.t.p., how many molecules of hydrogen gas are present in 2.5 L of the gas?

Ans: 6.7209×10^{22} molecules

Video explanation: <https://youtu.be/VcMcvzu04qE>

48. If 20 moles of nitrogen gas reacts with 33 moles of hydrogen gas to produce ammonia. What is the limiting reagent?

Ans: nitrogen

Video explanation: <https://youtu.be/kWJ53miggUA>

49. 1 mole of a, 2 moles of b and 3 moles of c are put into a container. What is the mole fraction of c?

Ans: $\frac{1}{2}$

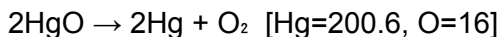
Video explanation: <https://youtu.be/aJCBJvmyDp0>

50. What volume of CO_2 is produced at s.t.p when 500 mL of 0.1 M Na_2CO_3 is reacted with excess HCl ?

Ans: 1.12 L

Video explanation: <https://youtu.be/FF2kGulnrYU>

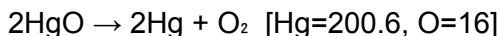
51. How many moles of HgO is needed to produce 120 g of oxygen in the equation



Ans: 7.5 moles of HgO

Video explanation: <https://youtu.be/SMnaod7-S1I>

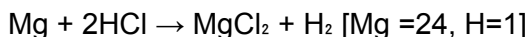
52. How many grams of Hg are produced from the reaction of 2.5 moles of HgO in the equation



Ans: 501.5 g

Video explanation: https://youtu.be/bMF1ys0_QvM

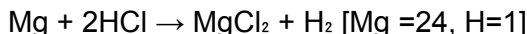
53. Calculate the mass of HCl that completely reacts with 3 moles of Mg in the equation



Ans: 219 g

Video explanation: <https://youtu.be/7G-0-reczSo>

54. If 5 moles of HCl reacts with 3 moles of Mg, how many grams of H₂ are produced?



Ans: 5 g of H₂

Video explanation: <https://youtu.be/SK1Mi4335Bk>

55. Calculate the molarity of 45 g of glucose in 250 mL of water.

Ans: 1 M

Video explanation: <https://youtu.be/KaDRMT0HFjQ>

56. Calculate the molarity of 54 g of LiCl in 300 mL of water [Li = 6.94, Cl = 35.5]

Ans: 4.24 M

Video explanation: <https://youtu.be/KnG9r2g5XyU>

57. Calculate the molarity of 10 mL of 6M HCl solution diluted to 150 mL with water.

Ans: 0.4 M

Video explanation: <https://youtu.be/pFs7I7AqINl>

58. Calculate the mass of solute in 250 mL of 0.75 M AgNO₃ solution [Ag = 108, N = 14, O = 16]

Ans: 31.875 g

Video explanation: https://youtu.be/i2FED_nbJPK

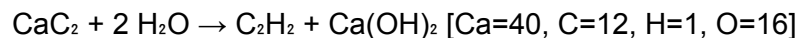
59. What mass of ethyne is produced when 3.5 moles of CaC_2 are completely reacted with water?



Ans: 91 g

Video explanation: <https://youtu.be/QQJyvf7FbLM>

60. How many grams of water are needed to completely react with 6 moles of CaC_2



Ans: 216 g

Video explanation: <https://youtu.be/CJ2JyJdM7vo>

61. What mass of sodium sulfate is needed to prepare 1.5 L of 0.55 M sodium sulfate solution?

Ans: 117.15 g

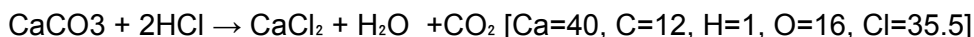
Video explanation: https://youtu.be/_ahLRBKe_Tl

62. What mass of glucose is needed to prepare 25 mL 1.75 M glucose solution?

Ans: 7.9 g

Video explanation: <https://youtu.be/4By5aWkeYMW>

63. If 750 mg of CaCO_3 is added to 20 mL of 0.5 M HCl, what mass of CO_2 is produced?



Ans: 0.22 g

Video explanation: <https://youtu.be/qerMT8yOmcM>

64. Calculate the molar mass of the acid if 0.35 g of the acid is dissolved in water then the solution is titrated using 11 mL of 0.55 M NaOH.

Ans: 57.85 g/mol

Video explanation: <https://youtu.be/uJqucyagmh8>

65. What is the molarity of 30 mL of 6 M ethanoic acid which is then diluted with 2500 mL water?

Ans: 0.071 M

Video explanation: <https://youtu.be/JN89FsmdLD0>

66. How many grams of NaHCO_3 would completely react with 27 mL of 6 M sulphuric acid?

Ans: 27.216 g

Video explanation: <https://youtu.be/FZ93UiRIDvk>

67. 1 L of air contains 0.28 g of oxygen. If John inhales 0.50 L of air per breath and 20 breaths per minute. How many grams of oxygen does he inhale per hour?

Ans: 168 grams of oxygen per hour

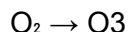
Video explanation: <https://youtu.be/4d2DZtYLJgl>

68. 10 mL of 0.55 M solution of $\text{Pb}(\text{NO}_3)_2$ is mixed with 9 mL of 0.15 M solution of Na_2SO_4 . What is the mass of PbSO_4 precipitated? [Pb=207, N=14, O=16, Na=23, S=32]

Ans: 0.4 g

Video explanation: <https://youtu.be/Hl0WmrQ0Obo>

69. How many grams of ozone can be produced from 5 g of oxygen based on the reaction below?



Ans: 5 g

Video explanation: <https://youtu.be/mh9xw-aaids>

70. 25 mL ethanoic acid solution requires 40 mL of 0.500 M KOH solution for complete titration. What is the molarity of ethanoic acid solution?

Ans: 0.8 M

Video explanation: <https://youtu.be/GvUozD1t-E4>

71. 40 g of metal is added to water in a cylinder. The volume of water in the cylinder increases by 4.5 mL. Calculate the density of the metal.

Ans: 8.89 g/mL

Video explanation: <https://youtu.be/5xUP8UQOqAQ>

72. If the combustion of 90 g of propene C_3H_6 produces 81% yield, what is the actual

amount of CO_2 produced? $\text{C}_3\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

Ans: 229.114 g

Video explanation: <https://youtu.be/V3h1kqVIG2U>

73. How much glucose (mm= 180.16 g/mol) do you need weigh to out to prepare 0.04 L of 0.016 M solution?

Ans: 0.1153 g

Video explanation: https://youtu.be/aymcC5np_Go

74. How much CaCl_2 (solid) is needed to make 10 mL of 0.25% (w/v) solution?

Ans: 0.025 g

Video explanation: <https://youtu.be/88VZhhdhQVCQ>

75. How many grams of K_2CO_3 ($M_r = 138.2 \text{ g/mol}$) are needed to make 600 mL of a 3 M solution?

Ans: 248.76 g

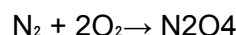
Video explanation: <https://youtu.be/qg9PD9XqTH8>

76. In a titration, 10 mL of 5 mM NaOH standard were needed to neutralise 50 mL of an unknown solution of acetic acid (CH_3COOH). What is the molarity of the ethanoic acid?

Ans: 0.001 M

Video explanation: <https://youtu.be/EgnUvzGGhZI>

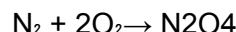
77. If 15 g of N_2O_4 was produced, how many moles of O_2 were required?



Ans: 0.326 moles

Video explanation: <https://youtu.be/XCxuoDCZhLI>

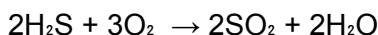
78. If 4×10^3 moles of oxygen reacted, how many grams of N_2 were required?



Ans: 56000 g

Video explanation: <https://youtu.be/CzKYH9zEEedM>

79. How many grams of oxygen gas is consumed to produce 4.6 mol SO_2 ?



Ans: 220.8 g

Video explanation: <https://youtu.be/JCwGTIzHu2M>

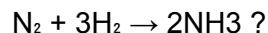
80. How many grams of $PbCr_2O_7$ is precipitated from the combination of 25 mL of 0.15 M $Pb(NO_3)_2$ and 15 mL of 0.10 M $K_2Cr_2O_7$ solutions?



Ans: 0.6345 g

Video explanation: <https://youtu.be/ghErEnXVWL0>

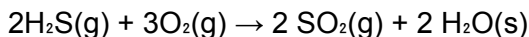
81. What volume of H_2 is needed to produce 35 L of ammonia at STP?



Ans: 52.5 L

Video explanation: <https://youtu.be/zAX0ts0ng74>

82. How many atoms of oxygen are needed to produce 18 g of ice?



Ans: 1.8066E24

Video explanation: https://youtu.be/h_cPCHYpdTg

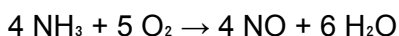
83. Determine the mass of oxygen produced when 70 g of $\text{Cu(NO}_3)_2$ decomposes.



Ans: 5.97 g

Video explanation: <https://youtu.be/uF9PRibyNjo>

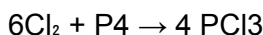
84. Determine the limiting reagent when 3 moles of NH_3 are reacted with 96 g of O_2



Ans: Oxygen

Video explanation: <https://youtu.be/4nj56mgg8Zo>

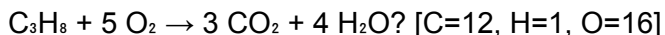
85. What mass of PCl_3 is formed from the reaction of 13 g of Cl_2 with 30 g of P_4 based on the reaction below?



Ans: 16.775 g

Video explanation: <https://youtu.be/YCmKqliOrrY>

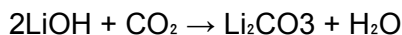
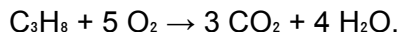
86. Will 80 g of oxygen be enough for the complete combustion of 35 g of C_3H_8 ?



Ans: Not enough

Video explanation: <https://youtu.be/wa8lzgp8BRA>

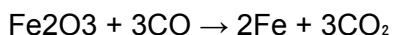
87. 35 g of C_3H_8 was completely combusted producing CO_2 (g). What mass of LiOH is needed to absorb all the CO_2 ?



Ans: 114.25 g

Video explanation: <https://youtu.be/XXqUbycpUjo>

88. In the experiment, 330 g of Fe is obtained. Calculate the percentage yield if we used all 500 g of iron oxide.



Ans: 94.28 %

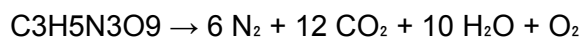
Video explanation: https://youtu.be/Af_GSKg8JxU

89. The percentage yield for the production of NH_3 is 70%. Calculate the mass of NH_3 obtained when 35 g of nitrogen reacts with 5 g of hydrogen gas.

Ans: 19.83 g

Video explanation: <https://youtu.be/JalBQflvdg>

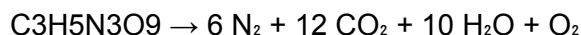
90. What amount of oxygen (in grams) is expected from the decomposition of 100 g of nitroglycerin ($\text{C}_3\text{H}_5\text{N}_3\text{O}_9$)?



Ans: 3.52 g

Video explanation: https://youtu.be/zRsS_ddD12I

91. Contrary to expectation, only 6.5 g of oxygen is produced from the the decomposition of 200 g of nitroglycerin ($\text{C}_3\text{H}_5\text{N}_3\text{O}_9$). Calculate the percentage yield?



Ans: 92.32%

Video explanation: <https://youtu.be/3Zu1KWeTR08>

92. Calculate the mass of oxygen formed when 20 g of KClO_3 decomposes based on the reaction below $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$

Ans: 7.836 g

Video explanation: <https://youtu.be/EU8XzKVf0I0>

93. How many grams of hydrogen are left over when 15 g of nitrogen reacts with 8 g of hydrogen to produce ammonia?

Ans: 4.78 g

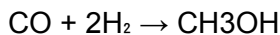
Video explanation: <https://youtu.be/kbfObqQEWjU>

94. How many grams of hydrogen gas react completely with 3.87×10^{23} particles of nitrogen in the reaction below? $\text{N}_2 + 3 \text{H}_2 \rightarrow 2 \text{NH}_3$

Ans: 3.855 g

Video explanation: https://youtu.be/_xMRXly9Pfc

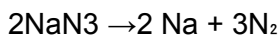
95. At STP, what volume of H_2 is needed to react completely with 5.56×10^{23} molecules of CO?



Ans: 41.36 L

Video explanation: https://youtu.be/1mqrdBP_hE4

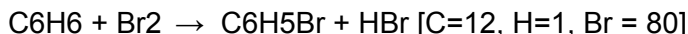
96. Calculate the mass of NaN_3 required to yield 50 L of N_2 (density = 1.25 g/L).



Ans: 96.72 g

Video explanation: <https://youtu.be/H1s7ta2MJCQ>

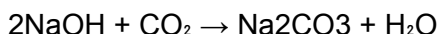
97. Calculate the theoretical yield of $\text{C}_6\text{H}_5\text{Br}$ when 15 g of C_6H_6 reacts with 35 g of Br_2 ?



Ans: 30.17 g

Video explanation: <https://youtu.be/yJVz0UPQVDw>

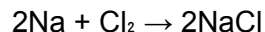
98. What mass of Na_2CO_3 is produced when 1.02 moles of NaOH reacts with 1 mol of CO_2 .



Ans: 54.06 g

Video explanation: <https://youtu.be/EkY1xWivXiE>

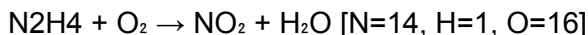
99. Salut Einstein! How many formula units of NaCl are produced from the reaction of 25 g of Na with 2 moles of chlorine gas?



Ans: 6.54×10^{23} formula units

Video explanation: <https://youtu.be/pAte9ZbVewA>

100. How many grams of oxygen would react completely with 165 g of hydrazine (N_2H_4) based on the reaction below?



Ans: 495 g of oxygen

Video explanation: <https://youtu.be/5bVChL4Zk4Y>

Hi! I'm happy you got this practice quiz. I also really hope you improve greatly on your Stoichiometry. Remember...practice leads to improvement! You can reach out to us at copychemistry22@gmail.com or help@copychemistry.com

Most people say we are the **fastest** when it comes to replies and **providing help**.

Instagram https://www.instagram.com/copy_chemistry

Facebook <https://www.facebook.com/copychemistry>

TikTok https://www.tiktok.com/@copy_chemistry

Threads https://www.threads.net/@copy_chemistry

Twitter https://x.com/copy_chemistry

Telegram <https://t.me/copychemistry>

YouTube https://youtube.com/@copy_chemistry

Email copychemistry22@gmail.com, help@copychemistry.com

Need Private Tutoring? Schedule now!: <https://tinyurl.com/mtkph26k>

Need Guidance With Assignment? Upload it here!: <https://tinyurl.com/4cf9emzc>

Find out more about us at www.copychemistry.com

